

Physics 2 - Homework 8 Solutions

Problem (A):

The molecular weight of glucose ($C_6H_{12}O_6$) is about 180 *g/mole* so

$$\frac{1 \text{ mole}}{0.18 \text{ kg}} \times \frac{1000 \text{ kg}}{1 \text{ tonne}} \times \frac{674 \text{ kcal}}{\text{mole}} = 3.7 \times 10^6 \text{ kcal}$$

Chapter 5:

10. a)

$$14 \text{ km}^2 \times \left(\frac{1000 \text{ m}}{1 \text{ km}}\right)^2 \times 12 \text{ m} \times \frac{1}{6 \text{ hr} \times 3600 \text{ s/hr}} = 7,778 \text{ m}^3/\text{s}$$

b)

$$\frac{7,778 \text{ m}^3/\text{s}}{7 \text{ m/s}} = 1,111 \text{ m}^2$$

11.

$$\frac{1000 \text{ lb}}{\text{person} \cdot \text{yr}} \times 2.9 \times 10^8 \text{ people} \times \frac{4300 \text{ Btu}}{\text{lb}} = 1.25 \times 10^{15} \text{ Btu} = 1.25 \text{ QBtu}$$

Comparing this to the U.S. energy use in 1996, we have $\frac{1.25}{94} = 1.3\%$.

Chapter 6:

1. a) ^{235}U contains 235 grams per mole, so:

$$\frac{1000 \text{ g}}{235 \text{ g/mole}} \times 6.02 \times 10^{23} \text{ atom/mole} \times 1.98 \times 10^6 \text{ eV/atom} \times 1.6 \times 10^{-19} \text{ J/eV} = 8.1 \times 10^{13} \text{ J}$$

b) $0.7\% \times 8.1 \times 10^{13} = 5.7 \times 10^{11} \text{ J}$

3. Table 2.7 on pg. 52 says there are 250×10^9 tonnes of coal reserves in the United States. In the front cover of the text the energy content of coal is given to be 2.81×10^{10} J/ton, where 1 ton = 0.9 tonne. Hence,

$$250 \times 10^9 \text{ tonnes} \times \frac{1 \text{ ton}}{0.9 \text{ tonnes}} \times \frac{2.81 \times 10^{10} \text{ J}}{\text{ton}} = 7.8 \times 10^{21} \text{ J}$$

From pg. 183, the total of the 2 lowest price classes for uranium is 3×10^6 tonnes. In the front cover the energy content is given as 8.28×10^{10} J/g, so

$$3 \times 10^6 \text{ tonnes} \times \frac{1000 \text{ kg}}{\text{tonne}} \times 0.007 \times 8.28 \times 10^{13} \text{ J/kg} = 1.7 \times 10^{21} \text{ J}$$

There is approx. 4.6 times less energy resource in uranium-235 compared to coal.

9. a) From Prob. 1 above, 1 tonne of U-235 has the equivalent of 8.1×10^{16} J, so the number of tonnes required, at 33% efficiency, is:

$$\frac{10^{12} \text{ W} \times 3.15 \times 10^7 \text{ s}}{(8.1 \times 10^{16}) \text{ J/tonne} \times 0.33} = 1,167 \text{ tonne}$$

b) $\frac{1167}{.007} = 1.67 \times 10^5 \text{ tonne}$

14. The number of half-lives in 48 days is $\frac{48}{8} = 6$, so the amount remaining is: $20 \text{ Ci} \times (\frac{1}{2})^6 = 3.125 \times 10^{-2} \text{ Ci}$.